Date Planned : __ / __ / __	Daily Tutorial Sheet-6	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level-2	Exact Duration :_____

76. Carbon is in the lowest oxidation state in :

(A) CH_4 **(B)** CCl_4 **(C)** CF_4 **(D)** CO_2

77. The oxidation number of carbon in CH_2Cl_2 is :

(A) 0 **(B)** +2 **(C)** –2 **(D)** +4

78. Sulphur has highest oxidation state in

(A) SO_2 **(B)** H_2SO_4 **(C)** $Na_2S_2O_3$ **(D)** $Na_2S_4O_6$

79. In which of the following compounds the oxidation number of carbon is maximum :

(A) HCHO **(B)** $CHCl_3$ **(C)** CH_3OH **(D)** $C_{12}H_{22}O_{11}$

80. An element which never has a positive oxidation number in any of its compounds is :

(A) Boron **(B)** Oxygen **(C)** Chlorine **(D)** Fluorine

81. The oxidation number of hydrogen in MH_2 is : (where M ≡ bivalent metal)

(A) +1 **(B)** –1 **(C)** +2 **(D)** –2

82. The most common oxidation state of an element is –2. The number of electrons present in its outermost shell is :

(A) 4 **(B)** 2 **(C)** 6 **(D)** 8

83. Which of the following reactions involves oxidation-reduction ?

(A) $NaBr + HCl \longrightarrow NaCl + HBr$ **(B)** $HBr + AgNO_3 \longrightarrow AgBr + HNO_3$

(C) $H_2 + Br_2 \longrightarrow 2HBr$ **(D)** $2NaOH + H_2SO_4 \longrightarrow Na_2SO_4 + 2H_2O$

84. The compound which can not act both as oxidising as well as reducing agent is :

(A) SO_2 **(B)** MnO_2 **(C)** Al_2O_3 **(D)** CrO

***85.** H_2S acts only as a reducing agent while SO_2 can act both as a reducing and oxidizing agent because :

(A) S in H_2S has –2 oxidation state

(B) S in SO_2 has oxidation state +4

(C) Hydrogen in H_2S is more positive than oxygen

(D) Oxygen is more negative in SO_2